EXPERIMENT # 14

NAME:_____

TABLE SALT FROM SODA ASH
REPORT FORM

DATA:

Mass of beaker _____ g

Mass of beaker and sample (Na_2CO_3) _____ g

Mass of sample (Na_2CO_3) _____ g

Mass of beaker and residue (NaCl) _____ g

Mass of residue (NaCl) _____ g

Volume of conc. HCl added _____ mL

CALCULATIONS:

Molar mass of Na_2CO_3 _____ g/mol

Number of moles of Na_2CO_3 added _____ mol

Show calculations below:

Concentration of HCl (aq) _____M

Number of moles of HCl added _____ mol

Show calculations below:

Write a balanced chemical equation for this chemical reaction. Include state designations for all reactants and products.

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Assuming Na_2CO_3 is the limiting reactant, how many grams of NaCl can be produced from the amount of Na_2CO_3 used in this experiment? ___ g
Show calculations below:

Assuming HCl is the limiting reactant, how many gran of NaCl can be produced from the amount of HCl usec this experiment? ___ g
Show calculations below:

Which reactant is limiting? _____

Which reactant is excess? _____

Actual yield of NaCl _____ g

Theoretical yield of NaCl _____ g

Percent yield of NaCl _____ %
Show calculations below:

Why must the first heating be done under the fume hood rather than at your bench?

_____

_____