

IONIC & COVALENT BONDING**Exit Ticket**

1. Classify each compound below as *ionic* or *covalent*:

a) PCl_3 _____

b) Rb_2O _____

c) CaCl_2 _____

d) BrF _____

2. Write Lewis structure for each ionic compound shown below:

a) CaCl_2

b) K_2O

3. Write Lewis structure for each covalent compound shown below:

a) NF_3

b) N_2O (oxygen is terminal)

c) H_2CO (carbon is central)

4. Write Lewis structure for each molecule shown below. Include resonance structures if necessary.



5. Draw Lewis structures for each molecule shown below, and predict the molecular shape of each:



6. Shown below is the structure arrangement for acetic acid. Complete the Lewis structure and determine the shape and bond angle for the three atoms marked a, b and c.

