TEST 1 REVIEW

1. Round the following measurements to 3 significant figures:
a) 0.0485624 m
b) 83000000 s
c) 1842700 g
2. Convert the following units:
a) 536 mg to kg
b) 253 mi to $\mathrm{cm}(1 \mathrm{mi}=1.6093 \mathrm{~km})$
c) $120 \mathrm{~km} / \mathrm{h}$ to $\mathrm{miles} / \mathrm{min}$
3. In each pair below, select the higher temperature.
a) $10^{\circ} \mathrm{C}$ or $40^{\circ} \mathrm{F}$
b) $\quad 300 \mathrm{~K}$ or $90^{\circ} \mathrm{F}$
4. Read each Celsius thermometer below with the correct number of digits:


B
$\qquad$
$\qquad$
5. Which diagram below illustrates heterogeneous mixture? $\qquad$

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6. Shown below are the specific heats of several substances. If equal amounts of each substance are heated at the same temperature,

| a) Which substance will have the highest temperature? |
| :--- | :---: | :---: |
| Briefly explain. | Substance | Specific heat |
| :---: |
| $\left(\mathbf{J} / \mathbf{g}^{\circ} \mathbf{C}\right)$ |
| Copper |
| Silver |
|  |
| Titanium |

a) Which substance will have the lowest temperature? Briefly explain.
7. How many joules of heat are required to raise the temperature of 975 g of aluminum from $0.0^{\circ} \mathrm{C}$ to $50.0^{\circ} \mathrm{C}$ ? The specific heat of aluminum is $0.900 \mathrm{~J} / \mathrm{g}^{\circ} \mathrm{C}$.
8. 25.0 kJ of heat are added to a $500.0-\mathrm{g}$ bar of iron metal at $25^{\circ} \mathrm{C}$. What is the final temperature of the iron bar? Specific heat of iron is $0.452 \mathrm{~J} / \mathrm{g}^{\circ} \mathrm{C}$.
9. What mass (in grams) of a $12.0 \%$ sugar water solution is needed to supply 20.0 g of sugar?
10. The price of 1 lb of potato is $\$ 1.75$. If all the potatoes sold at the store in one day bring in $\$ 1420$, how many kg of potato did the grocery shoppers buy?
11. Celeste's diet restricts her intake of protein to 24 g per day. If she eats an 8.0 -oz burger that is $15.0 \%$ protein, has she exceeded her protein limit for the day? If yes, how many ounces of a burger would be allowed for Celeste?

