

LOS ANGELES MISSION COLLEGE-SPRING 2020
CHEMISTRY 51-SECTION 10929 & 110947 & 18821

Lecture: TTh 4:35 – 6:40 pm ; Room: CMS-028

Lab (10947) T 6:50 -10:00 pm; Room: CMS-201

Lab (18821) Th 6:50 -10:00 pm; Room: CMS-203

INSTRUCTOR (LEC & LAB): Gayane Godjoian

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CLASS DESCRIPTION: Chem 51 is an introductory class in general chemistry and is designed for students in the following majors: Nursing, Allied Health Sciences; Dietetics, Physical Therapy, Food Science & Environmental & Occupational Health. This course may also be taken to satisfy the Physical Science requirement for General Education. Chemistry 51 at LA Mission College is equivalent to Chemistry 103 or Chemistry 105 at CSUN.

PREREQUISITE: Mathematics 115 (Elementary Algebra) or 123B (Elementary and Intermediate Algebra II) with a grade of “C” or better, or appropriate Math placement results.

REQUIRED MATERIALS:

- **Textbook:** “General, Organic, and Biological Chemistry”, by Timberlake, Custom Published for LAMC. (ISBN: 978-1-323-87902-3)
 - A copy of the textbook is available on Reserve in the Library.
- **Lab Manual:** No lab manual required for this class. Laboratory experiment directions are available on class website at www.profpaz.com.
- **Lab Notebook:** This is a **quadrille paper, hard cover “Comp Book”**, available in the L.A.Mission College Bookstore and in the C.S.U.N. Bookstore. You must have the Laboratory Notebook by the second class meeting. You are required to report all laboratory work in your Laboratory Notebook (See Appendix II for the proper use of the Laboratory Notebook) During the Laboratory Activities you are not permitted to take notes on any kind of loose paper or any notebook, other than the Laboratory Notebook. You may not perform an experiment if you do not have your Laboratory Notebook with you.
- **Periodic Table of Elements:** This is available in the LAMC bookstore and on my website. You must have one Periodic Table with you during all class sessions.
- **Scientific Calculator:** Need not to be an expensive type, but it must perform the following operations: Multiplication, Division, Addition, Subtraction, square root, 1/x, and log. You are required to have your calculator with you during all class sessions (both lectures and labs).
- **Safety Goggles:** Unless specifically instructed otherwise by your instructor, you must wear safety goggles during laboratory work. Safety goggles are available for purchase in the LAMC Bookstore. You are required to have your safety goggles by the second class session. You may keep your goggles locked in your laboratory locker.
- Failure to wear goggles when directed by the instructor is grounds for dismissal from the laboratory.
- **Notebook:** A 3-ring binder is recommended for keeping class notes organized.

**STUDENT
LEARNING
OUTCOMES:**

1. Conceptualize, model and explain chemical processes qualitatively at the molecular level.
2. Extract appropriate information, analyze and synthesize experimental results to reach correct conclusions.
3. Perform laboratory techniques safely and accurately and maintain a laboratory notebook according to standard scientific guidelines.

ATTENDANCE:

- CHEMISTRY IS A DEMANDING SUBJECT!
- YOU CANNOT AFFORD TO BE ABSENT IF YOU WISH TO DO WELL IN THIS COURSE.
- THERE IS NO MAKE-UP FOR MISSED LABORATORY WORK.
- College regulations state that a student may be excluded from a course following accumulation absences equal to a week of course work.

**COURSE
EVALUTATION:**

Your final grade in class is composed of the following:

<u>Assignment</u>	<u>Percentage</u>
Quizzes	10%
Exit Tickets	10%
Exams (3)	30%
Final Exam	15%
Lab Exams (2)	15%
Lab Reports	20%
Total	100%

**GRADING
SCALE:**

The final grades cutoffs are as follows:

A	90% - 100%
B	80% - 90%
C	65% - 80%
D	55% - 65%
F	Below 55%

NOTES:

- **No make up** exams are given for students being absent on the day of the exam.
- If serious and compelling reasons prevent the student from being present on the day of one of the exam, the instructor should be informed **IN ADVANCE** for possible arrangements.
- Maximum of one make-up exam and one make-up quiz per semester is allowed.

LABORATORY WORK

- During laboratory work two students will share the contents of the same laboratory locker.
- Both students are jointly responsible for the contents of their shared locker.
- The majority (not all) of the experiments are performed in pairs.
- **For every experiment, each student,**
 - 1. will take active part in the work,**
 - 2. report his/her data individually,**
 - 3. do his/her own calculations,**
 - 4. turn in an individual lab report for grading purposes and**
 - 5. will be assigned an individual grade for every activity.**
- Laboratory Reports are due one week after the class period in which the experiments have been performed (this is to allow working students to meet the deadline).
- **Late reports are subject to a penalty of 10% per class time.**
- Once the instructor has returned the graded lab reports to the class, lab reports for that particular experiment are no longer accepted for grading.
- In order to work efficiently and meet the required deadline for turning in the lab reports, **you must come** to the laboratory well prepared.
- **This means:**
 - 1. Read carefully (several times, if needed) the Experiment you will perform (both Principles and Procedure) prior to coming to the lab.**
 - 2. Think about what will be doing and plan ahead.**
 - 3. Prepare your Laboratory Notebook in advance (Purpose of the Experiment and the appropriate Data Tables may be prepared in your Laboratory Notebook in advance).**

THERE IS NO MAKE-UP LABORATORY WORK

INSTRUCTIONS FOR LABORATORY NOTEBOOK

- Each student must have a **quadrille ruled, sewn** Laboratory Notebook in which to record data and observations, do calculations, and analyze results of the lab work.
- The Lab Notebook must be brought with you to every lab session and all data and observations must be recorded **directly into the Notebook** (nowhere else) **and in ink** (no pencil). Laboratory records are legal documents in industry and research. They are required to support patent applications or to resolve disputes or originality of research .
- You will write only on the **right hand pages**. The left-hand pages are reserved for calculations and notes that do not belong on the right hand page.
- Begin with a **TITLE PAGE** State the course, section number, semester, the instructor's name, your name and your locker number.
- The second page is an **INDEX**. As you do each experiment, list it by title and enter the numbers of the pages containing text for it. Leave a second page for continuation of the Index. At the bottom of the second index page, give the **complete bibliographic information** for the laboratory text used. (Title, author, publisher, date.) When you do this you can cite a reference simply by "Text"; otherwise you must cite the complete reference each time.
- The remainder of the **right-hand pages** in the Notebook should be **numbered sequentially in the upper right corner of the page**.

The **FORMAT** of the pages for each lab experiment is as follows:

TITLE:	Here you enter the title of experiment.	Page Number
PURPOSE:	Write a short statement (one or two sentences, in your own words) of the purpose or the goal of the experiment.	
PROCEDURE:	Cite a reference to the appropriate text(s). Any changes made by the instructor may be noted on the left-hand side of the page.	
DATA/OBSERVATIONS:	Prepare a data table in which you will record the measurements you make in the lab. The lab Report Form often will provide a good format, but it is wise to check with the instructor about the amount of space to be allowed when observations, rather than measurements, are to be recorded. Be careful to indicate units wherever appropriate.	
RESULTS:	This presents, in table form, the final answers to any required calculations. All work (i.e., set-ups for all calculations) must be shown on the left-hand page .	
CONCLUSIONS:	Essentially, your conclusions should answer the Purpose or the Goal of the Experiment. Write a few words of conclusion, indicating any experimental errors and their effects on your results. Also state whether or not you achieved the purpose of the experiment.	

- As you work, enter your Data/Observations **in ink**. If you make an error or repeat an exercise, **DO NOT ERASE ANYTHING**. You may draw a line through the offending information and then enter the new value (It may be necessary to do this on the left-hand page, if there is no room on the right-hand page.)
- If the entire page is in error, simply draw a diagonal line through the page and fold the page in half vertically.
- **NEVER, NEVER, TEAR OUT A PAGE** (other pages will fall out as well).
- **BE PREPARED TO SHOW YOUR NOTEBOOK TO YOUR INSTRUCTOR AT ANY TIME!**
- Additional Information about the proper usage of the Laboratory Notebook is found in Appendix II of the Laboratory Manual used for this course (“Everyday Chemistry” by Maria Fenyes, Los Angeles Mission College, Fall 96)

TENTATIVE LECTURE SCHEDULE

Week	Date	Chapter	Lecture Topic	Text Reference
1	Feb. 11	1	Introduction to class Scientific Method; Key Math Concepts	----- 1.2-1.4
	Feb. 13	1 2	Scientific Notation SI Units; Significant Figures	1.5 2.1-2.3
2	Feb. 18	2	Unit Conversions; Density	2.4-2.7
	Feb. 20 Feb. 23	2	Unit Conversions; Density Last day to drop without a "W"(online only)	2.4-2.7
3	Feb. 25	3	Classification of Matter; Energy & Heat	3.1-3.6
	Feb. 27	3	Changes of State; Review for Test 1	3.7
4	March 3	-----	Test 1 (Chapters 1-3)	-----
	March 5	4	Periodic Table; Atomic Theory	4.1-4.3
5	March 10	4	Atomic Structure; Isotopes; Modern Atomic Theory	4.4-4.5
	March 12	4	Electron Configuration; Periodic Trends	4.7-4.8
6	March 17	6	Ionic Compounds – Naming and Writing Formulas	6.1-6.3
	March 19	6	Covalent Compounds – Naming and Writing Formulas	6.4-6.5
7	March 24	6	Molecular Shapes & Polarity	6.6-6.8
	March 26	7	Types of Chemical Reactions/Balancing Equations	7.1-7.2
8	March 31	----	Cesar Chavez-Campus Closed	-----
	April 2	-----	Test 2 (Chapters 4 & 6)	-----
9	April 14	7	Redox Reactions; Concept of Mole	7.3-7.5
	April 16	7	Calculations involving mole concept	7.6-7.7
10	April 21	7	Stoichiometry; Limiting Reactants / Percent Yield	7.8-7.9
	April 23	7	Energy and Chemical Reactions	7.10
11	April 28	-----	Review for Test 3	-----
	April 30	-----	Test 3 (Chapters 7)	-----
12	May 5	9	Solutions and Solubility	9.1-9.2
	May 7 May 10	9	Solubility Last Day to drop classes with a "W" (online only)	9.3
13	May 12	9	Net Ionic Equations	Notes
	May 14	9	Solution Concentrations & Properties	9.4-9.6
14	May 19	11	Acids & Bases; Ionization of Water	11.1-11.5
	May 21	11	pH Calculations; Reactions of Acids & Bases; Buffers	11.6-11.8
15	May 26	8	Gases and Their Properties; Gas Laws	8.1-8.5
	May 28	8	Gases in Chemical Reactions; Gas Mixtures	8.6-8.8
16	June 2 (5:30-7:30 pm)	-----	Final Exam (Chapters 8, 9 & 11)	-----

LABORATORY SCHEDULE

Week	Date	Exp. #	Activity	Points
1	T, Feb. 11	----	Introduction to Lab; Safety Video; Check-In	----
	Th, Feb. 13	----	Introduction to Lab; Safety Video; Check-In	----
2	T, Feb. 18	1	Safety Contract & Test Due & Measurements	10
	Th, Feb. 20	1	Safety Contract & Test Due & Measurements	10
3	T, Feb. 25	2	Density	10
	Th, Feb. 27	2	Density	10
4	T, Mar. 3	3	Qualitative Separation of Mixtures	10
	Th, Mar. 5	3	Qualitative Separation of Mixtures	10
5	T, Mar. 10	4	Quantitative Separation of Mixtures	10
	Th, Mar. 12	4	Quantitative Separation of Mixtures	10
6	T, Mar. 17	6	Specific Heat of a Metal	10
	Th, Mar. 19	6	Specific Heat of a Metal	10
7	T, Mar. 24	-----	Lab Exam I (Exp 1-4)	----
	Th, Mar. 26	-----	Lab Exam I (Exp 1-4)	----
8	T, Mar. 31	----	Cesar Chavez-Campus Closed	-----
	Th, Apr. 2	H/O	Nomenclature Worksheet and Writing & Balancing Equations	20
9	T, Apr. 14	10	Combination & Decomposition Reactions	20
	Th, Apr. 16	10	Combination & Decomposition Reactions	20
10	T, Apr. 21	11 & 12	Single and Double Replacement Reactions	20
	Th, Apr. 23	11 & 12	Single and Double Replacement Reactions	20
11	T, Apr. 28	9	Percentage of Copper in Malachite	10
	Th, Apr. 30	9	Percentage of Copper in Malachite	10
12	T, May 5	H/O	Net Ionic Equations and Electrolytes & Nonelectrolytes	20
	Th, May 7	H/O	Net Ionic Equations and Electrolytes & Nonelectrolytes	20
13	T, May 12	H/O	Properties of Acids & Bases	10
	Th, May 14	H/O	Properties of Acids & Bases	10
14	T, May 19	H/O	Charles's Law	10
	Th, May 21	H/O	Charles's Law	10
15	T, May 26	----	Lab Final Exam (Exp remainder of experiments) (You may use your lab notebook)	-----
	Th, May 28	-----	Lab Final Exam (Exp remainder of experiments) (You may use your lab notebook)	-----