Chemistry 51 Name:_____

EXPERIMENT # 12
REPORT FORM

Part A:

Evidence of Reaction	***Write a balanced molecular equation with state designation.***
1.	$NaCl\ (aq) + KNO_3\ (aq) \rightarrow$
2.	$NaCl\ (aq) + AgNO_3\ (aq) \rightarrow$
3.	$Na_2CO_3\ (aq) + HCl\ (aq) \rightarrow$
4.	$NaOH\ (aq) + HCl\ (aq) \rightarrow$
5.	$BaCl_2\ (aq) + H_2SO_4\ (aq) \rightarrow$
6.	$NH_4OH\ (aq) + H_2SO_4\ (aq) \rightarrow$
7.	$CuSO_4\ (aq) + Zn(NO_3)_2\ (aq) \rightarrow$
8.	$Na_2CO_3\ (aq) + CaCl_2\ (aq) \rightarrow$

9.	$CuSO_4$ (aq) + NH_4Cl (aq) →
10.	NaOH (aq) + HNO_3 (aq) →
11.	$FeCl_3$ (aq) + NH_4OH (aq) →
12.	Na_2SO_3 (aq) + HCl (aq) →

Answer the questions below:

1. Which reactions above can be classified as a precipitation reaction?

   _____ (indicate reaction numbers)

2. Which reactions above can be classified as a neutralization reaction?

   _____ (indicate reaction numbers)

Part B:

For each precipitation reaction in Part A, write:

I. Balanced molecular equation with state designations.
II. Balanced total ionic equation. (Be sure to include charges on ions)
III. Balanced net ionic equation. (Be sure to include charges on ions)

Use spaces below, as needed, to write your equations.

Molecular equation:
Total ionic equation:
Net ionic equation:
Molecular equation:
Total ionic equation:
Net ionic equation:
Molecular equation:
Total ionic equation:
Net ionic equation:
Molecular equation:
Total ionic equation:
Net ionic equation:
Molecular equation:
Total ionic equation:
Net ionic equation: