

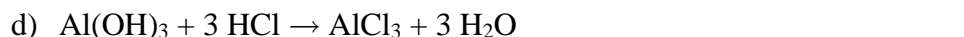
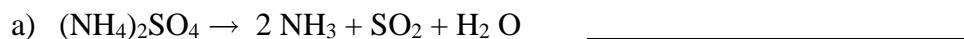
REVIEW QUESTIONS

Chapter 7

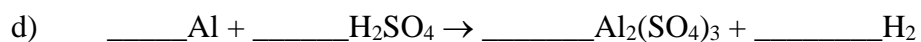
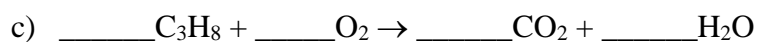
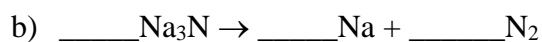
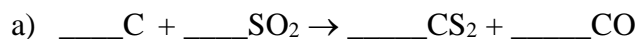
Calculate each of the following quantities:

1. Number of moles in 112 g of aspirin,  $\text{C}_9\text{H}_8\text{O}_4$
2. Mass of 3.82 moles of silver acetate,  $\text{AgC}_2\text{H}_3\text{O}_2$
3. Number of molecules in 1.75 moles of  $\text{CO}_2$
4. Number of molecules in 20.0 g of  $\text{CH}_4$

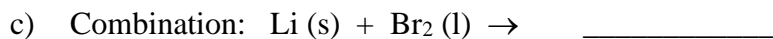
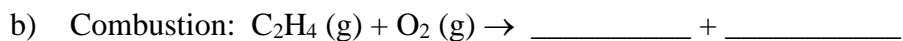
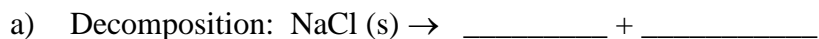
5. Classify the type of each of the following reactions:



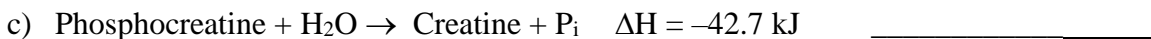
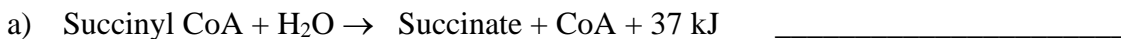
6. Balance each of the equations shown below:



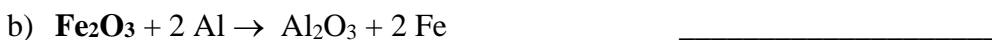
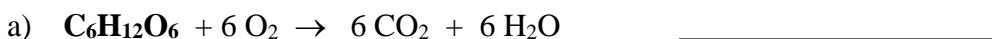
7. Predict the products and write a balanced equation for each of the following:



8. Each of the reactions below occur in the cells of the body. Identify each as endothermic or exothermic:



9. In each reaction below, identify the **bold type** substance as oxidized or reduced:



10. Chromium and oxygen combine to form chromium (III) oxide.

a) Write a balanced equation for this reaction.

b) How many moles of O<sub>2</sub> react with 4.50 mol of Cr?

c) How many grams of chromium (III) oxide are formed when 24.8g of Cr react?