Chemistry 51 Name:_____

ELE3CTROLYTES & EQUIVALENTS

Exit Ticket 11

1. Briefly provide definition and an example for each of the following terms:

   a) Strong electrolyte:_____

   _____

   _____

   b) Weak electrolyte:_____

   _____

   _____

   c) Non-electrolyte:_____

   _____

   _____

2. Determine the number of equivalents in each of the following:

   a) 0.30 mol of SO_4^{2-} : _____

   b) 5.0 mol of Al^{3}: _____

   c) 15 g of Ca^{2+} : _____

3. An IV solution contains 4.0 mEq/L of Ca^{2+}. How many g of Ca^{2+} are in 3.0 L of this solution?

4. Classify each of the following as **strong electrolyte (SE), weak electrolyte (WE)** or **non-electrolytes (NE)**:

   a) acetic acid: _____

   b) methanol (CH_3OH): _____

   c) potassium chloride (KCl): _____

5. Indicate whether each of the following compounds dissolves in water to give ions, molecules, or both.

   A) both
   B) ions
   C) molecules

   a) NaCl, a strong electrolyte: _____

   b) HI, a strong electrolyte: _____

   c) CH_3CH_2OH, a nonelectrolyte: _____

   d) KNO_3, a strong electrolyte: _____

   e) glucose, a nonelectrolyte: _____

   f) H_2CO_3, a weak electrolyte: _____