

**SOLUTIONS****Exit Ticket 9**

Use the solubility diagrams on the back or in your textbook to answer the questions below:

1. Identify each salt below as soluble or insoluble:

a)  $\text{CuCO}_3$  \_\_\_\_\_

c)  $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2$  \_\_\_\_\_

b)  $(\text{NH}_4)_2\text{S}$  \_\_\_\_\_

d)  $\text{AgI}$  \_\_\_\_\_

2. Pair each cation on the left with an anion on the right that will form a soluble compound, and write a formula. Use each anion only once.

$\text{Na}^+$

$\text{NO}_3^-$

$\text{Sr}^{2+}$

$\text{SO}_4^{2-}$

$\text{Co}^{2+}$

$\text{S}^{2-}$

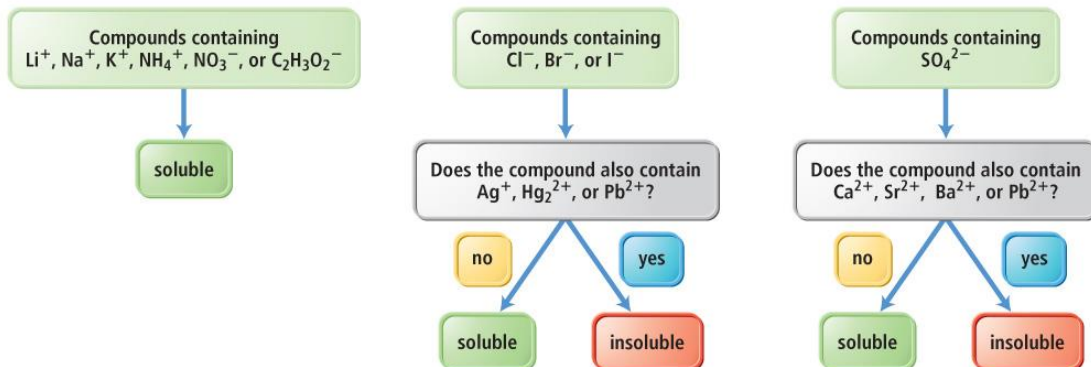
$\text{Pb}^{2+}$

$\text{CO}_3^{2-}$

3. Move any misplaced compounds to the correct column:

<b>Soluble</b>	<b>Insoluble</b>
LiOH	CaCl <sub>2</sub>
Na <sub>2</sub> CO <sub>3</sub>	Cu(OH) <sub>2</sub>
AgCl	Ca(C <sub>2</sub> H <sub>3</sub> O) <sub>2</sub>
K <sub>3</sub> PO <sub>4</sub>	SrSO <sub>4</sub>
PbBr <sub>2</sub>	CuI <sub>2</sub>
CoCO <sub>3</sub>	PbI <sub>2</sub>

### Compounds that are mostly soluble



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