pH Calculations

1. Complete the missing information in the table below:

$[H_3O^+]$	$[OH^-]$	pH	pOH	Acidic or Basic?
1.0×10^{-12} M				
		5.56		
	1.0×10^{-5}			
			10.85	
		8.40		
	6.8×10^{-11}			
3.2×10^{-3}				
			4.68	

2. Use the information in the table above to complete the following sentences:

   a) As the $[H_3O^+]$ of solution increases, the solution becomes more _____

   b) As the $[OH^-]$ of solution increases, the solution becomes more _____

   c) As the pH of solution increases, the solution becomes more _____

   d) As the pOH of solution increases, the solution becomes more _____

   e) As the acidity of a solution increases, its pH _____

   f) As the basicity of a solution increases, its pH _____

   g) As the acidity of a solution increases, its pOH _____

   h) As the basicity of a solution increases, its pOH _____

ANSWER KEY