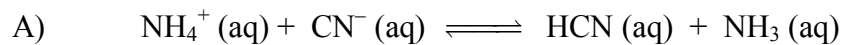
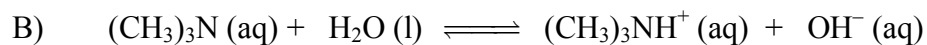


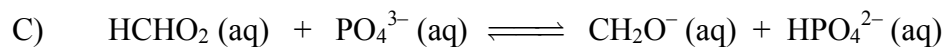
REVIEW QUESTIONS

Chapter 15

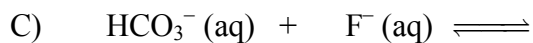
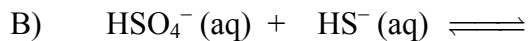
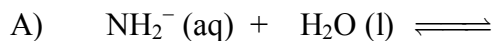
1. For each reaction below, identify the Brønsted-Lowry acid and base and their conjugates:







2. Predict the products of the following acid-base reactions, and use Table 16.2 in your text to determine whether the equilibrium lies to the right or left of the equation:



3. Calculate the $[\text{OH}^-]$ for each solution below, and indicate whether each solution is acidic, basic or neutral:

A) $[\text{H}^+] = 6.0 \times 10^{-6} \text{ M}$

B) $[\text{H}^+] = 2.5 \times 10^{-8} \text{ M}$

4. Calculate the $[\text{H}^+]$ and $[\text{OH}^-]$ for each of the following strong acids and bases:

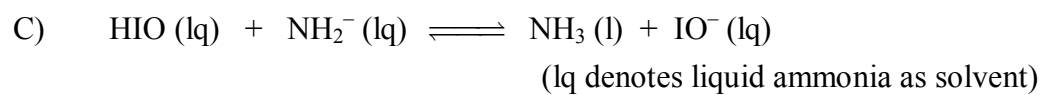
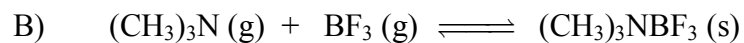
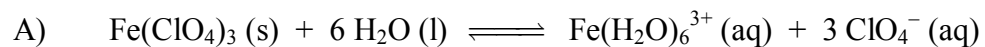
A) $1.8 \times 10^{-4} \text{ M HBr}$

B) $0.025 \text{ M H}_2\text{SO}_4$

C) $3.2 \times 10^{-3} \text{ M KOH}$

D) $0.0075 \text{ M Ca(OH)}_2$

5. Identify the Lewis acid and base in each of the following reactions:



6. For each pair, determine which is the stronger acid, and give a brief explanation:

