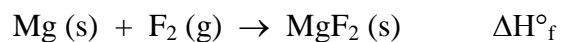


IONIC & COVALENT BONDING**Exit Ticket**

1. Use the following data to set up a Born-Haber cycle and calculate the enthalpy of formation (ΔH°_f) for magnesium fluoride. Be sure to write an equation that corresponds to each energy given, and show how addition of all steps in the cycle produces the equation for the formation of magnesium fluoride (shown below).



Lattice energy	-2913 kJ/mol
1 st IE of Mg	735 kJ/mol
2 nd IE of Mg	1445 kJ/mol
EA of F	-328 kJ/mol
Bond energy of F ₂	154 kJ/mol
Enthalpy of sublimation for Mg	150. kJ/mol