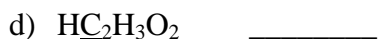
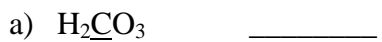
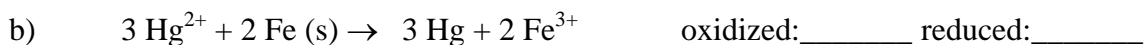


**Oxidation-Reduction Worksheet**

1. Determine the oxidation number of each element underlined:



2. Identify the species being oxidized and reduced in each of the following reactions:



3. Complete each sentence below by choosing the proper term:

a) When  $\text{ClO}_3^-$  reacts to form  $\text{ClO}_2$ , it acts as a (an) \_\_\_\_\_ agent.  
(oxidizing/reducing)b) When  $\text{H}_2\text{S}$  reacts to form  $\text{SO}_4^{2-}$ , it acts as a (an) \_\_\_\_\_ agent.  
(oxidizing/reducing)c) When  $\text{MnCl}_2$  reacts to form  $\text{MnO}_2$ , it acts as a (an) \_\_\_\_\_ agent.  
(oxidizing/reducing)d) When Zn metals reacts to form  $\text{ZnCl}_2$ , it acts as a (an) \_\_\_\_\_ agent.  
(oxidizing/reducing)

4. Balance each skeleton half-reaction below, and add to obtain the balanced overall equation:

